

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS

GCE Advanced Subsidiary Level and GCE Advanced Level

**MARK SCHEME for the May/June 2011 question paper
for the guidance of teachers**

9701 CHEMISTRY

9701/33

Paper 31 (Advanced Practical Skills 1),
maximum raw mark 40

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

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Question	Sections	Indicative material	Mark	
1 (a)	PDO Layout	I Volume given for rough titre and accurate titre details tabulated. <i>Minimum of 2 × 2 “boxes”.</i>	1	
	PDO Recording	II Appropriate headings and units for data given in weighing and accurate titration tables. <i>Acceptable headings:</i> <i>mass of tube + FA1;</i> <i>mass of tube + residue/mass of empty tube (mass of FA1 used);</i> <i>initial/final or 1st/2nd (burette) (reading)/(reading at start/finish);</i> <i>volume added/used/ titre; or wtte [not “difference”]</i> <i>Acceptable units are solidus: /cm³; brackets: (cm³); in words: volume in cubic centimeters, volume in cm³. Similarly for mass in g, etc</i> If units are not included in the heading every entry in the table must have the correct unit.	1	
	PDO Recording	III All accurate burette readings are given to the nearest 0.05 cm ³ . <i>Do not award this mark if:</i> <i>50(.00) is used as an initial burette reading;</i> <i>more than one final burette reading is 50.(00);</i> <i>any burette reading is greater than 50.(00)</i>	1	
	MMO Decision	IV Two uncorrected titres within 0.10 cm ³ <i>Do not allow the Rough even if ticked.</i> <i>Do not award this mark if having performed two titres within 0.1 cm³ a further titration is performed which is more than 0.10 cm³ from the closer of the initial two titres, unless a fourth titration, within 0.1 cm³ of any other has also been carried out.</i>	1	
Examiner rounds any burette readings to the nearest 0.05 cm ³ , checks subtractions and then selects the “best” titre using the hierarchy: <i>two identical; titres within 0.05 cm³; titres within 0.1 cm³; etc</i> to calculate mean (ignore any labelled rough). Examiner compares [corrected mean titre/corrected mass of FA 1] with Supervisor result. Calculate the ratios to 2 dp.				
	MMO Quality	Award V, VI and VII if $\delta \leq 0.05$ (cm ³ g ⁻¹) Award V and VI if $0.05 < \delta \leq 0.10$ Award V only if $0.10 < \delta \leq 0.20$ <i>If the “best” titres are ≥ 0.60 cm³ apart cancel one of the Q marks.</i>	1 1 1	[7]

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(b)	MMO Decision	Selects correctly subtracted accurate titre values within 0.20 cm^3 . Must use more than one value. If no calculation shown then titres must be indicated (e.g. with a tick) in the table	1	[2]
	PDO Display	<p>Correct mean from any values selected (may include rough) by candidate given to same decimal places as most precise burette reading recorded in the table.</p> <p><i>The third decimal place may be rounded to the nearest 0.05 cm^3.</i></p> <p><i>A mean of exactly .x25 or .x75 is allowed but the candidate may round up or down to the nearest 0.05 cm^3.</i></p> <p><i>If ALL burette readings are given to 1 decimal place then the mean may be given to 1 decimal place if numerically correct without rounding.</i></p> <p><i>Mean of 24.3 and 24.4 = 24.35 (✓)</i></p> <p><i>Mean of 24.3 and 24.4 = 24.4 (✗)</i></p> <p>If no working shown allow mean if value identical to that used by Examiner.</p>	1	
(c)	ACE Interpretation	I In part (i) $\{\text{titre from (b)}/1000\} \times 0.01(0)$ <i>If no working shown then answer must be correct.</i>	1	[5]
		II ans to (i) $\times 5$ and ans to (ii) $\times 10$ <i>with no additional steps</i>	1	
		III ans to (iii) $\times 55.8$ <i>If (iii) incorrect allow correct (ii) $\times 10 \times 55.8$</i>	1	
		IV correct (ans to (iii) $\times 55.8/\text{mass of FA 1}) \times 100$ to sf shown (<i>ecf</i> allowed from (iii)) (sf shown may come from (i) with no previous rounding) <i>If (iii) incorrect allow correct (ii) $\times 10 \times 55.8 \times 100/\text{mass FA 1}$</i> <i>(If choice of answer take the one in the answer space.)</i>	1	
	PDO Display	V 3 or 4 significant figures in final answers to all parts attempted (minimum three parts)	1	
(d)	ACE Interpretation	(i) Uncertainty either 1 or .5 in final place. If balance displays to 1 decimal place: error in balance reading is $\pm 0.05 \text{ g}$ or $\pm 0.1(0) \text{ g}$ If balance displays to 2 decimal places: error in balance reading is $\pm 0.005 \text{ g}$ or $\pm 0.01 \text{ g}$ If balance displays to 3 decimal places: error in balance reading is $\pm 0.0005 \text{ g}$ or $\pm 0.001 \text{ g}$	1	[2]
		(ii) $\{2 \times (\text{i})/\text{mass used}\} \times 100$ answer to 2, 3 or 4 sf	1	
				[Total: 16]

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2 (a)	PDO Display	I Tabulates mass of (empty) crucible, mass of crucible + FA 4 , mass of crucible + residue/ FA 4 after heating, mass of FA 4/hydrated magnesium sulfate, mass of water lost. <i>Do not award if mass of FA 4 or mass of water incorrect</i>	1	
	PDO Recording	II Records all weighings consistently to at least 1 decimal place [minimum three weighings].	1	
	MMO Decision	III Final weighings after reheating are within 0.05g or identical if masses recorded to 1 dp	1	
<p>Examiner to calculate [lowest mass of residue/mass of FA 4] of Supervisor and candidate to 2 dp. If two experiments carried out then use sum of masses of residues and sum of masses of FA 4 to calculate the ratio. <i>If the Supervisor's value is doubtful (higher than the majority of candidates) then check whether candidates are close to the expected value of 0.55 or use candidate average if majority in close agreement. (Contact team leader)</i></p>				
	MMO Quality	Award IV and V if $\delta \leq 0.05$ Award IV only if $0.05 < \delta \leq 0.10$	1 1	[5]
(b)	ACE Interpretation	Part (i) [mass of water lost/(7 × 18)] or [mass of water lost × 246.4/18] (allow $m(\text{H}_2\text{O}) \times 246.4/7$) <i>M_r H₂O must be 18</i> <i>Allow ecf for mass of water lost</i> <i>If two experiments carried out then mass of water may be taken from either or the mean.</i> <i>Allow mark if answer calculated correctly but working not shown.</i>	1	
		Part (i) [mass of water lost × 246.4/7 × 18] correctly evaluated to 3 sf [= mass of water lost × 1.956] (<i>Ignore part (ii)</i>) There are other chemically correct methods – mark accordingly.	1	[2]
(c)	ACE Improvements	Use a lid (for the early gentle heating) or larger mass (for smaller percentage error) or use (cool in) desiccator	1	[1]
			[Total: 8]	

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FA 5 is NaHCO ₃ (s); FA 6 is NH ₄ Br(s); FA 7 is H ₂ SO ₄ (aq)					
3	(a)	MMO Collection	On heating, steam or condensation or water vapour, misty vapour is noted or solid becomes powdery	1	[2]
		MMO Decisions	Tests for gas using limewater or in 3(d)	1	
(b)		PDO Layout	Presents results of tests in an unambiguous way <i>Minimum 4 × 2 boxes</i>	1	[6]
		MMO Collection	(No reaction with cold NaOH and) gas/ammonia/fumes produced (on heating) that turn(s) red litmus blue <i>Do not award if ppt reported with NaOH (CON)</i>	1	
			No reaction with ammonia and no reaction with barium chloride/nitrate	1	
			Cream ppt with silver nitrate that partially dissolves/is insoluble in aqueous ammonia	1	
		ACE Conclusion	FA 6 cation: ammonium/NH ₄ ⁺ <i>from some evidence and no CON obs</i> FA 6 anion: bromide/Br ⁻ <i>No ecf but can award Br⁻ from any mention of cream but ppt must be present or off-white ppt insoluble or partially soluble in NH₃.</i>	1 1	
(c)		MMO Collection	Ignore any observations after water added. Steamy/misty white/orange/red/red-brown (not brown) gas/vapour/ fumes/smoke produced or gas/vapour/fumes/smoke bleaches litmus (paper) or gas/vapour/fumes/smoke turns (potassium) dichromate (solution) from orange to green	1	[3]
		ACE Conclusion	(White) solid turns red/orange (not yellow, not brown, not solution, not ppt) Ignore “hot” FA 6 is oxidised/redox reaction/oxidation because Br ⁻ becomes Br ₂ /Br ₂ is produced or redox/reduction because H ₂ SO ₄ forms/becomes SO ₂ (with positive dichromate observation) or exothermic because tube becomes hot/heat given out.	1	
(d)		MMO Collection	Fizzing/effervescence/bubbling (occurs) (<i>not gas is produced</i>) If limewater test used here give second mark in (a).	1	[2]
			White ppt with lead nitrate and no reaction with silver nitrate	1	

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(e)	MMO Decision	barium chloride/nitrate followed by hydrochloric/nitric acid (not $\text{Ba}^{2+}(\text{aq})$, BaNO_3 , ...) (If H^+ already identified then “followed by hydrochloric/nitric acid” is not essential.)	1	[3]
	ACE Conclusion	FA 7 cation: protons/ H^+ if there is a positive observation with blue litmus paper/ K_2CrO_4 / Mg / Na_2CO_3	1	
		FA 7 anion: sulfate/ SO_4^{2-} Allow from minimum evidence of white ppt with $\text{Ba}^{2+}(\text{aq})$	1	
[Total: 16]				